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temperature, and their relation to energy, radiation, and physical properties of matter. The behavior of these quantities is governed by the four laws of thermodynamics which convey a quantitative description using measurable macroscopic physical quantities , but may be explained in terms of microscopic constituents by

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Thermodynamics - Wikipedia

Thermodynamics, science of the relationship between heat, work, temperature, and energy. In broad terms, thermodynamics deals with the transfer of energy from one place to another and from one form to another.

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The key concept is that heat is a form of energy corresponding to a definite amount of mechanical work.

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can do anything: move from one area to another, get atoms excited, and even increase energy. Did we say energy? That's what heat is.

Physics4Kids.com: Thermodynamics & Heat: Introduction

Basic physical laws govern how heat transfer for doing work takes place and

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Let us break the word thermodynamics into two words, thermo and dynamics.

'Thermo' stands for heat while

'dynamics' is used in connection with a mechanical motion which involves

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'work'. Therefore, Thermodynamics is the branch of physics that deals with the relationship between heat and other forms of energy.

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of fixed identity, boundaries may be fixed or movable, can transfer heat and work across boundary but not mass.

Force \times distance (work) System boundary Heat (Q) Electrical energy (work) System boundary.

THERMODYNAMICS: COURSE INTRODUCTION

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And that's pretty much what thermodynamics is, it's about, it's the study of heat and temperature, and how it relates to energy and work, and how different forms of energy can be transferred from one form to another.

And that's actually the heart of the first law of thermodynamics which we touched on on the introduction to energy

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The roles of heat transfer and internal

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energy change vary from process to process and affect how work is done by the system in that process. We will see that the first law of thermodynamics

explains that a change in the internal energy of a system comes from changes in heat or work.

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